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Verwey’s Hopping Transition Mechanism in relation to Dielectric studies of Zn & Sb substituted Cu ferrites

R. Dhanaraju a, M.K. Rajua, V. Brahmajirao b, S. Bangarraju c

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b Dept. Of Nano science and Technology, School of Biotechnology, MGNIRSA, c. Swaminathan Research Foundation,[DSRF], Gaganmahal, HYDERABAD-500029, Andhra Pradesh, INDIA
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Abstract: Hydroxyapatite (Ca 10 (PO 4) 6 (OH) 2; HAP) is a major mineral component of the calcified tissues, and it has various applications in medicine and dentistry.
Dependence of magnetic and structural properties of Ni$_{0.5}$M$_{0.5}$Fe$_2$O$_4$ (M=Co, Cu) nanoparticles synthesized by citrate precursor method on annealing temperature

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$^a$ Thermal Transport Laboratory, School of Chemical and Materials Engineering (SCME), National University of Sciences and Technology (NUST), H-12, Islamabad, Pakistan

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Study of the chelating/fuel agents influence on NiFe$_2$O$_4$ samples with potential catalytic properties

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http://dx.doi.org/10.1016/j.powtec.2013.03.033, How to Cite or Link Using DOI

Synthesis, characterization and magnetic properties of carbon nanotubes decorated with magnetic M\textsuperscript{II}Fe\textsubscript{2}O\textsubscript{4} nanoparticles

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Corresponding author. Tel.: +92 51 2077308; fax: +92 51 2077395.

Highlights: Magnetic carbon nanotubes were prepared by decorating with M\textsuperscript{II}Fe\textsubscript{2}O\textsubscript{4} nanoparticles. ► A simple microemulsion method was first time used for synthesis of M\textsuperscript{II}Fe\textsubscript{2}O\textsubscript{4}/CNTs. ► Carbon nanotubes were uniformly coated with large number of magnetic nanoparticles. ► M\textsuperscript{II}Fe\textsubscript{2}O\textsubscript{4}/CNTs nanocomposites show ferromagnetic behavior at room temperature.
Google Search: Thermal, XRD, and magnetization studies on ZnAl2O4 and NiAl2O4 spinels, synthesized by citrate precursor method and annealed at 450 and 650 C

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a. L.Lazau et al.Politechnica university of Victoriei, Romania

Cited No- 4- Thermal studies of some biologically active complexes containing 8-hydroxyquinolinate, DOI-10 1007/s10973-012-2511-3

a. Reena Sharma and neeraj Sharma, Dept. of Chemistry, Himachal University Shila, India
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Paper: Thermal, structural and magnetic studies on chromite spinel synthesized using citrate precursor method and annealed at 450 and 650°C

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Unlike ferrites, the chromite spinels have no B... Annealing has also been done at 650 °C in an attempt to obtain nano- metric particles of chromites (MCr2O4) (M = Cu, Ni... RK Singh, Patna women's College et al. ...

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b Department of Bioorganic Chemistry, University of Economics, 53-345 Wroclaw, Poland

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Synthesis, phonon and optical properties of nanosized CoCr2O4

M. Mączka, M. Ptak, M. Kurnatowska, J. Hanuza

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b Department of Bioorganic Chemistry, University of Economics, 53-345 Wroclaw, Poland

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Nanocrystalline Zn$_{0.5}$Ni$_{0.5}$Fe$_2$O$_4$
Preparation and kinetics of thermal process of precursor

Wenwei Wu · Yongqi Li · Kaiwen Zhou · Xuexiang Wu · Sen Liao · Qing Wang

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Abstract Zn$_{0.5}$Ni$_{0.5}$Fe$_2$(C$_2$O$_4$)$_2$·6H$_2$O was synthesized by solid-state reaction at low heat using ZnSO$_4$·7H$_2$O, NiSO$_4$·6H$_2$O, FeSO$_4$·7H$_2$O, and Na$_2$C$_2$O$_4$ as raw materials. The spinel Zn$_{0.5}$Ni$_{0.5}$Fe$_2$O$_4$ was obtained via calcining Zn$_{0.5}$Ni$_{0.5}$Fe$_2$(C$_2$O$_4$)$_2$·6H$_2$O above 773 K in air. The Zn$_{0.5}$Ni$_{0.5}$Fe$_2$(C$_2$O$_4$)$_2$·6H$_2$O and its calcined products were characterized by thermogravimetry and differential scanning calorimetry (TG/DSC), Fourier transform IR, X-ray powder diffraction (XRD), scanning electron microscopy (SEM), energy dispersive X-ray spectrometer (EDS), and vibratory sample magnetometer (VSM). The result showed that Zn$_{0.5}$Ni$_{0.5}$Fe$_2$O$_4$ obtained at 1073 K had a saturation magnetization of 86.7 emu g$^{-1}$. The thermal process of Zn$_{0.5}$Ni$_{0.5}$Fe$_2$(C$_2$O$_4$)$_2$·6H$_2$O experienced three steps, which involved the dehydration of the six crystal water molecules at first, and then decomposition of Zn$_{0.5}$Ni$_{0.5}$Fe$_2$(C$_2$O$_4$)$_2$ into Zn$_{0.5}$Ni$_{0.5}$Fe$_2$O$_4$ in air, and at last crystallization of Zn$_{0.5}$Ni$_{0.5}$Fe$_2$O$_4$. Based on KAS equation, and OFW equation, the values of the activation energies associated with the thermal process of Zn$_{0.5}$Ni$_{0.5}$Fe$_2$(C$_2$O$_4$)$_2$·6H$_2$O were determined to be 120.02 ± 23.93, and 259.76 ± 18.67 kJ mol$^{-1}$ for the first, and second thermal process steps, respectively. Dehydration of the six waters of Zn$_{0.5}$Ni$_{0.5}$Fe$_2$(C$_2$O$_4$)$_2$·6H$_2$O is multi-step reaction mechanisms. Decomposition of Zn$_{0.5}$Ni$_{0.5}$Fe$_2$(C$_2$O$_4$)$_2$ into Zn$_{0.5}$Ni$_{0.5}$Fe$_2$O$_4$ could be simple reaction mechanism, probable mechanism function integral form of thermal decomposition of Zn$_{0.5}$Ni$_{0.5}$Fe$_2$(C$_2$O$_4$)$_2$ is determined to be $g(x) = [-\ln(1 - x)]^2$.

Keywords Nanoparticles · Ferrites · Chemical synthesis · Non-isothermal kinetics · Thermal process

Introduction
Polycrystalline spinel ferrites have many unique properties, such as high electrical resistivity, high Curie temperature, large magnetocrystalline anisotropy, high coercivity, mechanical hardness, chemical stability, and temperature specific saturation magnetization, which make ferrites suitable for many applications in the field of high-density information storage, ferrofluids, catalysts, drug targeting, magnetic separation, magnetic resonance imaging, and gas sensor [1-6], etc. Within this group, Ni-Zn ferrites are very important soft magnetic materials. Its properties were highly dependent on the molar ratio of Ni to Zn, and crystallite diameter. Compared with other composition Ni$_x$Zn$_{1-x}$Fe$_2$O$_4$, Zn$_{0.5}$Ni$_{0.5}$Fe$_2$O$_4$ has higher specific saturation magnetizations [7].

To date, various methods have been developed to synthesize Ni$_x$Zn$_{1-x}$Fe$_2$O$_4$ with spinel structure, including co-precipitation [8, 9], citrate precursor method [10, 11], solid-state reaction at low heat [12], solid-state reaction at high temperature [13], sol-gel synthesis [13], glass-ceramic route [14], mechanical-chemical synthesis [15], molten salt method [16], refluxing method [17], reverse micelle method [18, 19], hydrothermal treatment [20], and combustion reaction [21], etc. In the synthesis of Zn$_{0.5}$Ni$_{0.5}$Fe$_2$O$_4$, it was found that crystallite diameter, morphology, and crystalline phases of Zn$_{0.5}$Ni$_{0.5}$Fe$_2$O$_4$ associated with its properties were highly dependent on the synthesis method and temperature. Such as Moualem-Bahout et al. [11] obtained globule-like...
The thermal behavior of some polymeric precursors used in CaAl\textsubscript{12}O\textsubscript{19} synthesis

I. Lazău · C. Păcurariu · R. Băbuţă

Abstract CaAl\textsubscript{12}O\textsubscript{19} was synthesised using three different precursors: (a) a polymeric type precursor resulted from the traditional Pechini method; (b) a polymeric type precursor resulted from the reaction between citric acid and calcium and aluminum nitrates; and (c) a polymeric type precursor resulted from the reaction between acryl acid and calcium an aluminum nitrates. The thermal behavior of the three precursors used in the CaAl\textsubscript{12}O\textsubscript{19} synthesis was monitored to underline the thermal effects associated to the CaAl\textsubscript{12}O\textsubscript{19} formation. Thermal analyses performed on precursors do not reveal clear differences regarding the thermal effects assigned to calcium aluminates formation, at temperatures over 800 °C. In contrast, thermal analysis of samples pre-fired at 200 °C, and especially at 600 °C, show clear differences between samples obtained in different ways. It is noted that in samples obtained from acrylic acid and nitrates, citric acid and nitrates, CA\textsubscript{16} is practically single phase after calcination at 1,200 °C. However, in the sample obtained from citric acid, ethylene glycol, and nitrates, calcined at 1,200 °C, CA\textsubscript{16} is present along with CA\textsubscript{2} and α-Al\textsubscript{2}O\textsubscript{3}.

Keywords Calcium aluminates · Polymeric precursor method · Thermal analysis

Introduction

Over time, the binary compounds of the CaO–Al\textsubscript{2}O\textsubscript{3} system were studied with interest for the valuable features owned by them. These compounds have been extensively analyzed due to their refractory and hydraulic properties and are also candidate materials for a wide range of technological applications because of their optical, electrical, thermal, and mechanical properties.

Calcium hexaaluminate CaAl\textsubscript{12}O\textsubscript{19} or CaO6Al\textsubscript{2}O\textsubscript{3} (CA\textsubscript{6}), known as bionite, has high refractory and a series of interesting electrical and optical properties. As a result, CA\textsubscript{6} synthesis is the subject of a large number of papers [1–5]. Since the formation of CA\textsubscript{6} through ceramic method runs difficult and requires very high temperatures (1,650 °C), the attention of researchers turned to various unconventional methods of synthesis, aiming the temperature decrease for the CA\textsubscript{6} synthesis.

Ianoş and his team [1] prepared single-phase CA\textsubscript{6} by low-temperature combustion synthesis, using a mixture of urea and β-alanine as fuel, no further ignition being necessary. Altay et al. [2] obtained CA\textsubscript{6} at 1,175 °C starting from a mixture of stoichiometric proportions of calcium and aluminum nitrates into 5 wt% aqueous solution of poly(vinyl alcohol). A new eco-friendly fabrication process for porous ceramics using hydraulic alumina and water was developed by Nagaoka et al. [3]; in this way they obtained CA\textsubscript{6} as single crystalline phase at 1,600 °C for 2 h. Using aluminum sulfate solutions and calcium nitrate as starting reagents, Singh and Sharma [4] obtained calcium hexaaluminate in the temperature range of 1,000–1,400 °C. Hexagonal CaAl\textsubscript{12}O\textsubscript{19} powders have been obtained at 1,200 °C for 2 h by a reverse micelle process [5].

Some calcium aluminates compounds were formed using as a raw material a hazardous waste from tertiary aluminum industry. The first aluminates formed was C\textsubscript{12}A\textsubscript{7} which at 838–848 °C had a transition to CA\textsubscript{2}, and then it was transformed to CA at 1,000–1,034 °C [6].
Thermal studies of some biologically active oxovanadium(IV) complexes containing 8-hydroxyquinolinate and hydroxamate ligands

Reena Sharma · Neeraj Sharma

SATAC-ACCT2011 Conference Special Chapter
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Abstract The thermal decomposition behaviours of oxovanadium(IV)hydroxamate complexes of composition [VO₂Q₂(OH)₂(HL)₂]⁻ [VO₂(polly)[Q]₂(OH)₂(HL)₂]⁻ [VO₂(polly)[Q]₂(OH)[5-Cl](CO)[CONH]O]⁻ (II), [VO₂(C₆H₅ON)[Q]₂(OH)[5-Cl](CO)[CONH]O]⁻ (III), and [VO₂(C₆H₅ON)[Q]₂(OH)[5-Cl](CO)[CONH]O]⁻ (IV) (where Q = C₆H₅NO⁻ 8-hydroxyquinolinate ion; HL₁⁻ = [C₆H₅NO(OH)H][CONH]O⁻ salicylhydroxamate ion; HL₂⁻ = [C₆H₅NO(OH)[5-Cl](CO)[CONH]O⁻ 5-chlorosalicylhydroxamate ion; n = 1 and 2), which are synthesised by the reactions of [VO₂(Q)₂]⁻ with predetermined molar ratios of potassium salicylhydroxamate and potassium 5-chlorosalicylhydroxamate in THF + MeOH solvent medium, have been studied by TG and DTA techniques. Thermograms indicate that complexes (I) and (III) undergo single-step decomposition, while complexes (II) and (IV) decompose in two steps to yield VO₂HL₁⁻,VO₂HL₂⁻ as the likely intermediate and VO₂ as the ultimate product of decomposition. The formation of VO₂ has been authenticated by IR and XRD studies. From the initial decomposition temperatures, the order of thermal stabilities for the complexes has been inferred as III > I > II > IV.

Keywords Oxovanadium(IV) complexes · Potassium salicylhydroxamate · Potassium 5-chlorosalicylhydroxamate · Thermal studies

Introduction

Hydroxamic acids and their derivatives, the weak organic acids with low toxicity of general formula R-CO-NH(OH) and R-CO-NR'OH have extensively been studied as bioligands forming chelate complexes with numerous metals [1, 2]. The utility of hydroxamic acids as colorimetric reagents for the separation and determination of metal ions, antimalarial and tumour inhibitor drugs, enzyme inhibitors, cell division and growth factors [3–5] has also gained enormous importance. Likewise, 8-hydroxyquinolinate and its derivatives constitute another group of biologically important ligands exhibiting predominantly antimicrobial activities [6–9]. These two groups of ligands have been of an enormous research interest in the coordination chemistry of vanadium owing to the role of vanadium complexes in nitrogen fixation, catalysis, design of molecular magnets, material science, as insulin mimetic, antitumour and antiamoebic agents [10–13]. The chemistry of oxovanadium(IV) and (V) ions: VO²⁺, VO³⁺, VO₄²⁻ and VO₂(O₃)⁺, in particular, because of their affinity towards a variety of ligands exhibiting diverse geometries around vanadium has attained phenomenal growth over the years. In view of the biological importance of vanadium on one hand and those of hydroxamate and 8-hydroxyquinolinate ligands on the other hand and in continuation of our interest on the synthesis of new oxovanadium(IV) complexes [14, 15], we investigated the potential of unexplored [VO₂(Q)₂]⁻ as precursor towards the synthesis of mixed ligand and quinolinate-free oxovanadium(IV) complexes using biologically important hydroxamate ligands viz. salicylhydroxamate and 5-chlorosalicylhydroxamate [16]. Owing to the versatility and considerable prominence of thermal methods in virtually all the branches of science and technology [17–20], it was imperative to gain an insight into the thermal behaviours of the newly synthesised complexes derived from the ligands (Fig. 1) attracted by the scattered reports in the literature on the thermal behaviours of vanadium complexes yielding V₂O₅, V₂O₃, and VO₂ as residual
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